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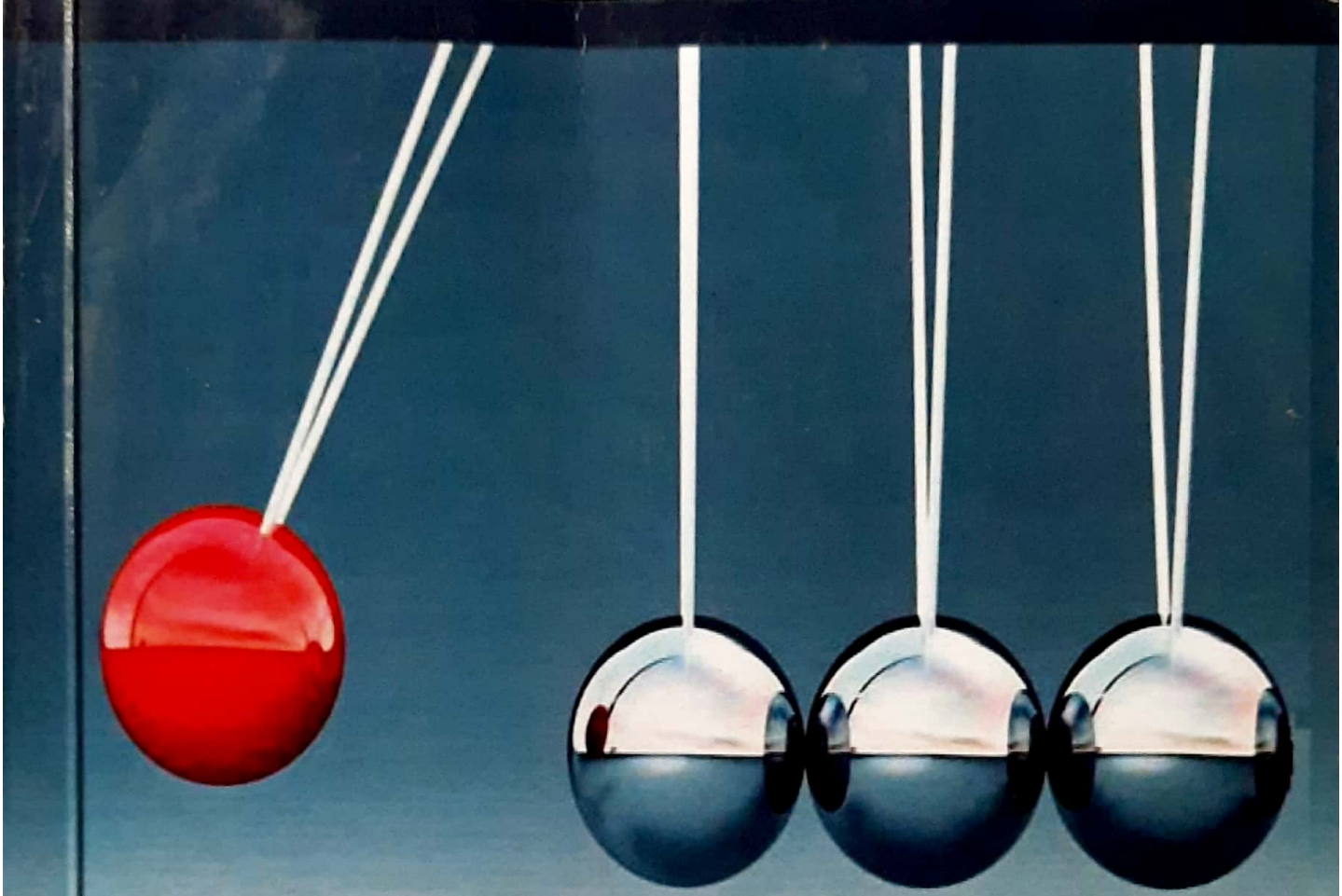

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For B.Sc. Students

Editor : Dr. C. M. Kale

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Senior Professor
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Dr. B. A. M. U. Aurangabad.

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(b) On the strength of its electrolyte
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(b) Higher than the battery voltage
(c) The same as the battery voltage
(d) None of the above

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(a) There is a rise in voltage
(b) Energy is absorbed by the cell
(c) Specific gravity of H_2SO_4 is increased
(d) All of the above

55]. A typical output of a solar cell is, - - -
(a) 0.1 V
(b) 0.26 V
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56]. Which of the following material is used in solar cells?
(a) Barium
(b) Silicon
(c) Silver
(d) Selenium

57]. In a lead acid cell, hydrogen is liberated at, - - -
(a) Positive plate
(b) Negative plate
(c) Both positive and negative plates
(d) None of the above

58]. Satellite power requirement is provided through, - - -
(a) Solar cells
(b) Dry cells
(c) Nickel Cadmium cells
(d) Lead acid batteries

59]. During the charging of a lead acid cell, - - -
(a) Its voltage increases (b) It gives out energy
(c) Its cathode becomes dark chocolate brown in Colour
(d) Specific gravity of H_2SO_4 decreases

60]. Active materials of a lead acid cell are, - - -
(a) Spongy lead (b) Lead peroxide
(c) Dilute H_2SO_4 **(d) All of the above**

61]. During charging the specific gravity of the electrolyte of a lead acid battery, - - -
(a) Increases (b) Decreases
(c) Remains the same (d) Becomes zero

62]. The capacity of a lead acid cell does not depend on its, - - -
(a) Temperature **(b) Rate of charge**
(c) Rate of discharge (d) Quantity of active material

63]. In a lead-acid cell dilute sulphuric acid approximately comprises the following, - - -
(a) One part H_2O, three parts H_2SO_4
(b) Two-part H_2O, two parts H_2SO_4
(c) Three parts H_2O, One part of H_2SO_4
(d) All H_2SO_4

64]. The watt-hour efficiency of a lead acid cell varies between, - - -
(a) 25 to 35% (b) 40 to 60%
(c) 70 to 80% (d) 90 to 95%

65]. The capacity of a lead acid cell depends on, - - -
(a) Amperes **(b) Ampere-hours**
(c) Watts (d) Watt-hours

66]. The capacity of a lead acid cell depends on, - - -
(a) Rate of discharge (b) Temperature
(c) Density of electrolyte **(d) All of the above**

67]. Level of electrolyte in a cell should be - - - the level of plates
(a) Below (b) Equal to
(c) Above (d) None of the above

68]. In a lead acid cell, lead is called as, - - -
(a) Positive active material
(b) Negative active material
(c) Passive material
(d) None of the above

69]. Electrolyte used in a lead acid cell is, - - -
(a) NaOH **(b) H_2SO_4**
(c) HCL (d) HNO_3

70]. The lead acid cell never be discharged beyond, - - -
(a) 1.8 V (b) 1.9 V
(c) 2 V (d) 2.1 V

71]. If a lead-acid cell is discharge below 1.8 V, the following will happen, - - -
(a) Capacity of cell will reduce
(b) Sulphation of plates will occur
(c) Internal resistance will increase
(d) All of the above

72]. In a lead acid battery the energy is stored in the form of, - - -
(a) Charged ions **(b) Chemical energy**
(c) Electrostatic energy (d) Electromagnetic energy

73]. The forbidden band exists in, - - -
(a) Semiconductor (b) Insulator
(c) Conductor **(d) Both a and b**

74]. No forbidden band exists between the valence band and the conduction band in a, - - -
(a) Conductor (b) Insulator
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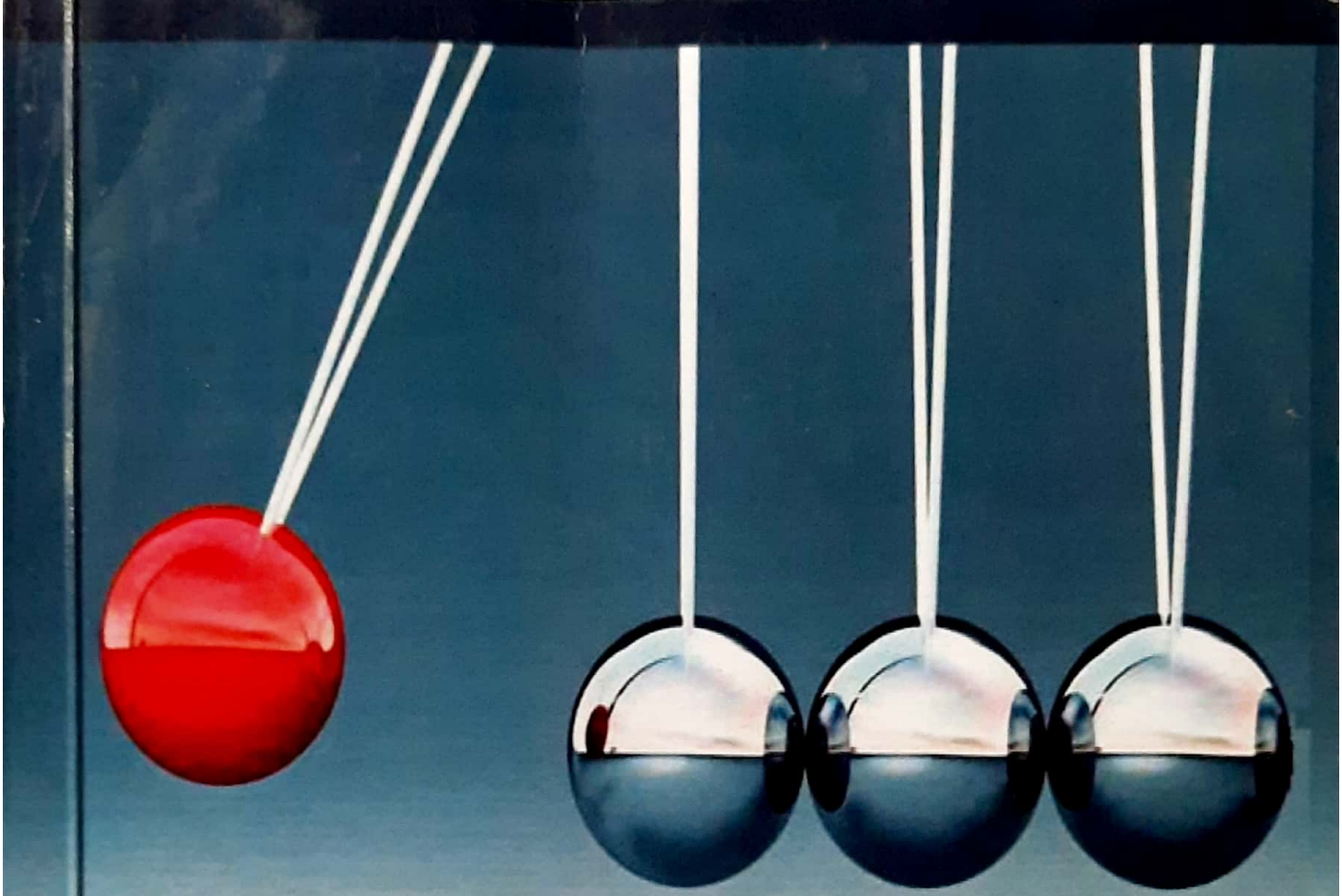
75]. The current flow through electrolyte is due to the movement of, - - -
(a) Ions (b) Holes
(c) Electrons (d) None of the above



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54]. It is noticed that during charging, - - -
(a) There is a rise in voltage
(b) Energy is absorbed by the cell
(c) Specific gravity of H_2SO_4 is increased
(d) All of the above

55]. A typical output of a solar cell is, - - -
(a) 0.1 V
(b) 0.26 V
(c) 1.1 V
(d) 2 V

56]. Which of the following material is used in solar cells?
(a) Barium
(b) Silicon
(c) Silver
(d) Selenium

57]. In a lead acid cell, hydrogen is liberated at, - - -
(a) Positive plate
(b) Negative plate
(c) Both positive and negative plates
(d) None of the above

58]. Satellite power requirement is provided through, - - -
(a) Solar cells
(b) Dry cells
(c) Nickel Cadmium cells
(d) Lead acid batteries

59]. During the charging of a lead acid cell, - - -
(a) Its voltage increases (b) It gives out energy
(c) Its cathode becomes dark chocolate brown in Colour
(d) Specific gravity of H_2SO_4 decreases

60]. Active materials of a lead acid cell are, - - -
(a) Spongy lead (b) Lead peroxide
(c) Dilute H_2SO_4 **(d) All of the above**

61]. During charging the specific gravity of the electrolyte of a lead acid battery, - - -
(a) Increases (b) Decreases
(c) Remains the same (d) Becomes zero

62]. The capacity of a lead acid cell does not depend on its, - - -
(a) Temperature **(b) Rate of charge**
(c) Rate of discharge (d) Quantity of active material

63]. In a lead-acid cell dilute sulphuric acid approximately comprises the following, - - -
(a) One part H_2O, three parts H_2SO_4
(b) Two-part H_2O, two parts H_2SO_4
(c) Three parts H_2O, One part of H_2SO_4
(d) All H_2SO_4

64]. The watt-hour efficiency of a lead acid cell varies between, - - -
(a) 25 to 35% (b) 40 to 60%
(c) 70 to 80% (d) 90 to 95%

65]. The capacity of a lead acid cell depends on, - - -
(a) Amperes **(b) Ampere-hours**
(c) Watts (d) Watt-hours

66]. The capacity of a lead acid cell depends on, - - -
(a) Rate of discharge (b) Temperature
(c) Density of electrolyte **(d) All of the above**

67]. Level of electrolyte in a cell should be - - - the level of plates
(a) Below (b) Equal to
(c) Above (d) None of the above

68]. In a lead acid cell, lead is called as, - - -
(a) Positive active material
(b) Negative active material
(c) Passive material
(d) None of the above

69]. Electrolyte used in a lead acid cell is, - - -
(a) NaOH **(b) H_2SO_4**
(c) HCL (d) HNO_3

70]. The lead acid cell never be discharged beyond, - - -
(a) 1.8 V (b) 1.9 V
(c) 2 V (d) 2.1 V

71]. If a lead-acid cell is discharge below 1.8 V, the following will happen, - - -
(a) Capacity of cell will reduce
(b) Sulphation of plates will occur
(c) Internal resistance will increase
(d) All of the above

72]. In a lead acid battery the energy is stored in the form of, - - -
(a) Charged ions **(b) Chemical energy**
(c) Electrostatic energy (d) Electromagnetic energy

73]. The forbidden band exists in, - - -
(a) Semiconductor (b) Insulator
(c) Conductor **(d) Both a and b**

74]. No forbidden band exists between the valence band and the conduction band in a, - - -
(a) Conductor (b) Insulator
(c) Semiconductor (d) None of these

75]. The current flow through electrolyte is due to the movement of, - - -
(a) Ions (b) Holes
(c) Electrons (d) None of the above

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